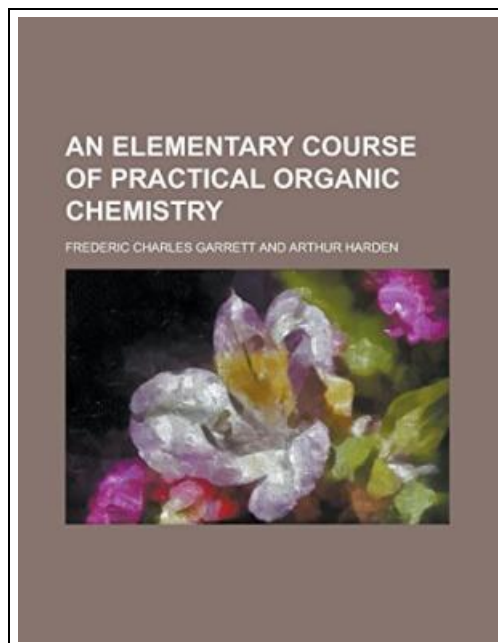


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*A high quality ebook as well as the typeface employed was exciting to read. It is actually loaded with wisdom and knowledge You wont sense monotony at at any moment of the time (that's what catalogues are for concerning when you request me).
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AN ELEMENTARY COURSE OF PRACTICAL ORGANIC CHEMISTRY



Rarebooksclub.com, United States, 2013. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1897 edition. Excerpt: . (4) Add a drop of concentrated sulphuric acid to a little pure aldehyde in a test-tube. A considerable evolution of heat occurs, and the aldehyde is converted into the polymeric paraldehyde, which boils at 124. (5) Warm a little aldehyde with aqueous caustic soda. A yellowish resinous mass is produced. (6) Dissolve a little aldehyde in water, add dilute sulphuric acid and a dilute solution of potassium permanganate. The latter is rapidly decolorised. (7) Make a very dilute solution of the colouring matter known as magenta, and add sulphurous acid until the liquid after shaking is colourless. A drop or two of aldehyde added to this solution restores the characteristic colour of magenta (Schiff's reaction). Chloral, $\text{CCl}_3\text{-CHO}$.--Chloral is trichlor-acetaldehyde, and reacts in most respects like acetaldehyde. Repeat the experiments, given under aldehyde, with a solution of chloral hydrate in water, and note the results. Decomposition of Chloral by Alkalis.--When treated with an alkali, chloral yields chloroform and a formate (hydrolysis): $\text{CCl}_3\text{-CHO} + \text{NaOH} = \text{CHCl}_3 + \text{HCOONa}$. Dissolve 6 grams of chloral hydrate in 5 grams of water, add 1 gram of caustic soda dissolved in 3 cc. of water, shake well, and allow to stand. Chloroform is precipitated and sinks to the bottom of the tube. Separate, dry, and if possible determine its boiling point. Acidify the aqueous layer with sulphuric acid, distil, and test the distillate for formic acid (p. 48). Preparation of Chloroform from Alcohol.--Chloroform may be prepared directly...



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